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## Biochar made from coffee ground for adsorption of crystal violet dye

### ABSTRACT

Biochar was prepared by anaerobic heating coffee ground at 800 °C for 3 hours. The prepared sample was used to adsorb crystal violet dye in an aqueous environment. The result showed that the biochar sample had good adsorption capacity for crystal violet dye, with the maximum adsorption capacity of 128.87 (mg/g). Thus, low-cost biochar prepared from coffee grounds has great potential to be used as an adsorbent material to remove crystal violet dye in an aqueous environment.

**Keywords:** Biochar, crystal violet, coffee ground, isotherm models, adsorption kinetics, wastewater treatment.

### 1. INTRODUCTION

Persistent organic dyes are often found in wastewater from the textile industry. A typical example is crystal violet dye (Cristal Violet – CV) (C<sub>25</sub>H<sub>30</sub>N<sub>3</sub>Cl), a strong color pollutant that hinders light penetration through water, causing harm to organisms. In addition, some reports have shown that CV can be dangerous to humans such as respiratory and kidney failure, skin and gastrointestinal irritation, cytotoxicity, and CV is even reported to be a carcinogen [1]. Therefore, removing CV in the aquatic environment is an urgent issue to create a safe living environment for humans and organisms. Recently, biochar has been used as a friendly, inexpensive adsorbent material to remove organic dyes in the aquatic environment [2]. Although the adsorption efficiency of biochar is lower than that of activated carbon, biochar is much cheaper due to the simple fabrication process from diverse biomass sources. Several studies have reported on the fabrication of biochar to remove organic dyes from agricultural biomass sources such as melia azadirach seeds [3], peanut shells [4], straw [5], walnut shells [6], date palm kernels [7], and cashew shells [8]. In this study, we fabricated biochar from coffee grounds using a simple thermal decomposition method, then the biochar samples

were used to adsorb CV in an aqueous medium. Different parameters of the adsorption experiment such as adsorption time and initial concentration of CV were investigated. The mechanism of CV adsorption by biochar was also evaluated based on the Langmuir and Freundlich isotherm adsorption models.

### 2. EXPERIMENT AND METHODS

#### 2.1. Process of making biochar from coffee ground

The coffee ground was thoroughly washed to remove impurities and dried at 80 °C in a drying oven for 24 hours to obtain dried coffee ground. Then, the coffee ground was anaerobically calcined at 800 °C for 3 hours. At the end of the calcination process, the biochar obtained was ground in a porcelain mortar and passed through a 60 µm sieve. The manufactured biochar sample was denoted as biochar-800 °C.

#### 2.2. Crystal violet adsorption experiment

The biochar samples were placed in Erlen flasks (250 mL capacity) containing 60 mL of crystal violet solution. The reaction flasks were sealed with ground glass stoppers and placed in a thermostatic shaking system (GFL 1083).

At the end of the reaction, centrifugation was performed to separate the solid fractions, and the solutions were used to measure the remaining CV concentration using a UV-VIS spectrophotometer. The adsorption efficiency of CV was calculated using the formula:  $R = \frac{C_0 - C}{C_0} \cdot 100\%$

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Where:

$C_0$  - initial concentration;  $C$  - remaining concentration after adsorption;  $R$  - removal (%).

### 2.2.1. Adsorption equilibrium time

Adsorption equilibrium time was determined by the experiments at the adsorption time of 0, 10, 20, 30, 40, 50, 60, 70, 80, 90, 100, 110, and 120 minutes. The remaining parameters such as the mass of adsorbent material is 0.06 g; CV concentration is 20 mg/L; stirring speed is 150 rpm.

### 2.2.2. Effect of CV concentration

The experiment on the effect of CV concentration on the adsorption process of biochar material was carried out by changing the initial concentration of CV to 20, 30, 40, 50, 60, 70, 80, 90, and 100 mg/L, respectively. The remaining parameters such as the mass of adsorbent material is 0.06 g; stirring speed is 150 rpm; and reaction time is determined as the time to reach adsorption equilibrium in the above experiment.

### 2.2.3. Adsorption isotherm equation

The adsorption isotherm equation was constructed based on the experimental results of using biochar to adsorb CV at different concentrations. The experimental data were calculated and fitted with the Langmuir and Freundlich isotherm adsorption models as shown in equations (1) and (2).

Langmuir isotherm model:

$$\frac{C_e}{Q_e} = \frac{C_e}{Q_m} + \frac{1}{K_L \cdot Q_m} \quad (1)$$

Freundlich isotherm adsorption model:

$$\ln Q_e = \ln K_F + \frac{1}{n} \cdot \ln C_e \quad (2)$$

The  $Q_e$  value is calculated as follows:

$$Q_e = \frac{(C_0 - C_e) \cdot V}{m} \quad (3)$$

Here,  $C_e$  (mg/L) and  $Q_e$  (mg/g) represent the concentration and adsorption capacity at equilibrium, respectively;  $Q_m$  (mg/g) is the maximum adsorption capacity;  $K_L$  is the Langmuir constant,  $K_F$  is the Freundlich constant,  $n$  is the Freundlich coefficient.

## 3. RESULTS AND DISCUSSION

### 3.1. CV adsorption standard curve

The standard curve shows the relationship between CV concentration and the maximum adsorption intensity constant Abs. The standard curve was determined by measuring the UV-Vis spectrum of CV solutions at different concentrations from 1 mg/L to 15 mg/L. The UV-Vis spectrum of CV solutions and the standard curve are shown in Fig. 1. The constructed standard curve was used to determine the concentration of CV solutions after being adsorbed by biochar.

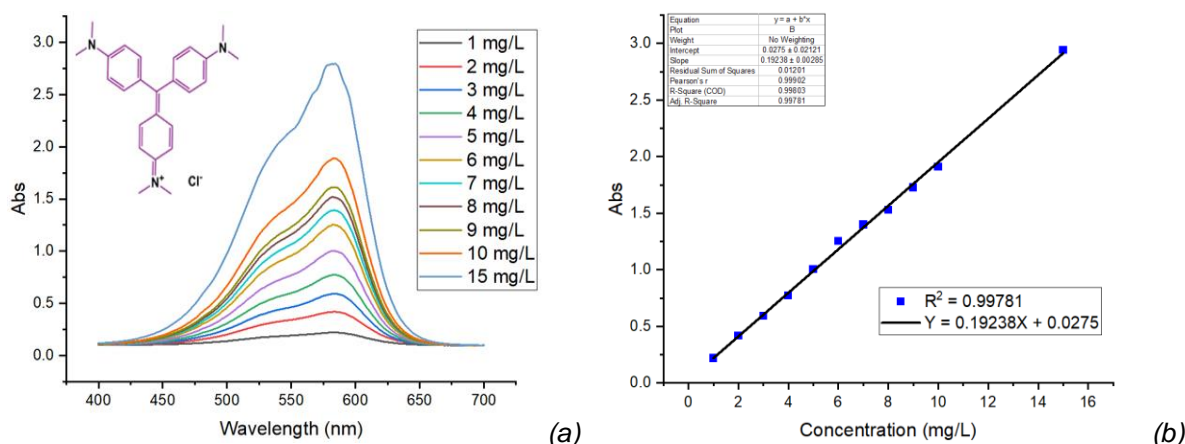


Figure 1. UV-Vis spectra of CV solutions at different concentrations (a) and adsorption standard curve (b)

### 3.2. Adsorption equilibrium time

Figure 2 shows the results of the investigation of the adsorption capacity of CV solution according to the reaction time of biochar. It was observed that the adsorption capacity of CV dye increased rapidly in the first 10 minutes, with an adsorption efficiency of 89.3%. Subsequently, the adsorption efficiency increased gradually, reaching 96.5% at 60 minutes, beyond which no significant increase was

observed. Thus, it can be seen that the adsorption equilibrium occurred within 60 minutes. After this time, the content of adsorbed CV remained almost unchanged because the active centers of the biochar adsorbent were completely occupied [9]. Thus, the CV adsorption capacity of biochar was quite fast in a short time of only 10 minutes, after which the adsorption rate gradually slowed down and reached equilibrium.

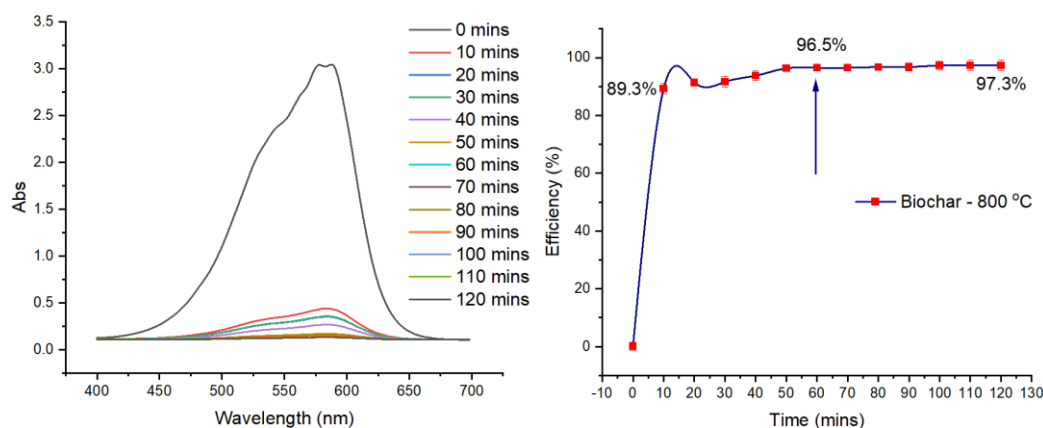


Figure 2. UV-Vis spectra of CV solution after adsorption by biochar at different times and the dependence of adsorption efficiency on time

### 3.3. Adsorption capacity of biochar with CV solution at different concentrations

Figure 3 presents the results of the investigation of the effect of the initial concentration of CV dye on the adsorption efficiency of biochar. The results show that as the initial concentration of CV dye increases, the adsorption capacity of the coal sample decreases. This is because the biochar adsorbent has enough active sites at low dye concentrations. After all, the vacant active sites on the adsorbent surface are filled with CV dye [9,10]. Surface saturation of the adsorbent

may occur when the concentration of the dye increases. Saturation occurs due to the limited surface area of the adsorbent, leading to a decrease in the removal efficiency of CV dye in the aqueous environment. Although the adsorption efficiency decreases as the initial concentration of CV increases, the adsorption efficiency of the biochar sample is quite high (over 90%) for all initial concentrations of the CV solution examined, demonstrating the adsorption efficiency of CV dye by biochar in this study.

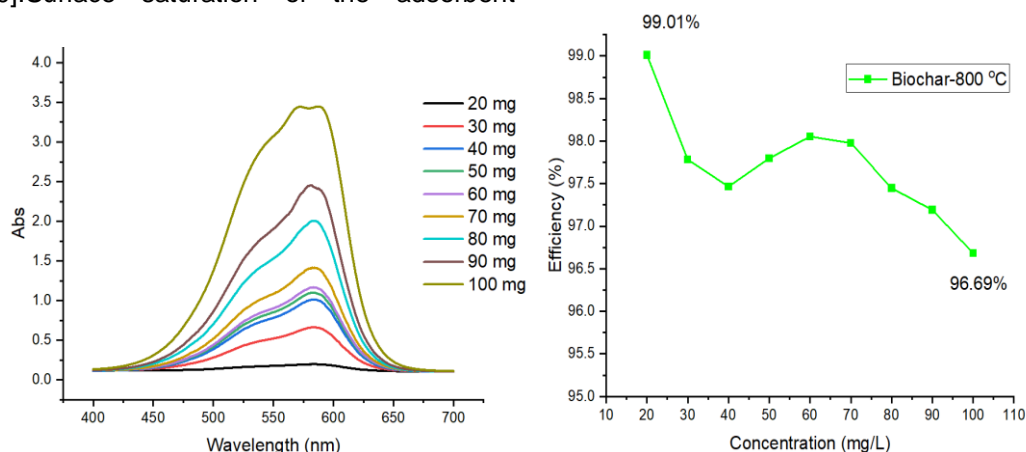


Figure 3. CV adsorption efficiency of biochar investigated at different concentrations

### 3.4. Isotherm adsorption mechanism

To determine the adsorption isotherm equations, experiments using TSH to adsorb CV crystal violet solution at different concentrations were performed as shown above. Table 1 presents the experimental and calculated data to construct the Langmuir and Freundlich isotherm adsorption models. From the calculated values, the adsorption isotherms were constructed using Origin software

as shown in Figure 4. Figure 4a shows the Langmuir isotherm adsorption model of the biochar sample. The results show that the biochar sample shows a relative agreement with the Langmuir isotherm adsorption model with a correlation coefficient value of  $R^2 = 0.75509$ .

The linear equation  $Y = AX + B$  according to the Langmuir isotherm adsorption model is  $Y = 0.00776X + 0.07342$ . From this equation, the

maximum adsorption capacity value  $Q_m$  (mg/g) of the biochar sample can be calculated according to the formula  $1/Q_m = A$  ( $A$  is the coefficient in the equation  $Y = AX + B$ ). The calculation results show that  $Q_m = 128.87$  (mg/g). The maximum adsorption capacity of the biochar sample in this study is relatively high, similar to the biochar samples reported in previous studies [7,8]. Figure 4b shows

the Freundlich isotherm adsorption model of the biochar sample. The results show that the correlation coefficient  $R_2$  is relatively high ( $R_2 = 0.89289$ ), indicating that the CV adsorption law of the biochar sample is more consistent with the Freundlich isotherm adsorption model [8,9]. The calculated parameters such as  $K_F = 17.782$  and  $n = 1.81$  are quite similar to previous studies [8,9].

Table 1. Data for building Langmuir and Freundlich adsorption isotherm models

$C_o$	20	30	40	50	60	70	80	90	100
$C_e$	0.886	3.314	5.128	5.580	5.928	7.218	10.471	12.998	17.094
$Q_e$	19.114	26.686	34.872	44.420	54.071	62.783	69.529	77.002	82.906
$C_e/Q_e$	0.046	0.124	0.147	0.126	0.110	0.115	0.151	0.169	0.206
$\ln C_e$	-0.121	1.198	1.635	1.719	1.780	1.977	2.349	2.565	2.839
$\ln Q_e$	2.950	3.284	3.552	3.794	3.990	4.140	4.242	4.344	4.418

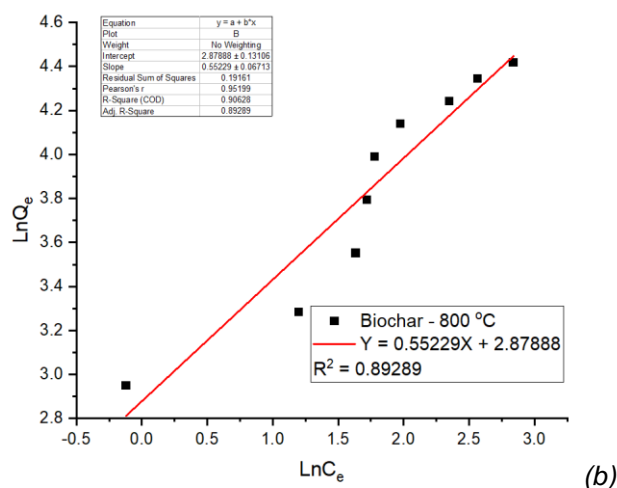
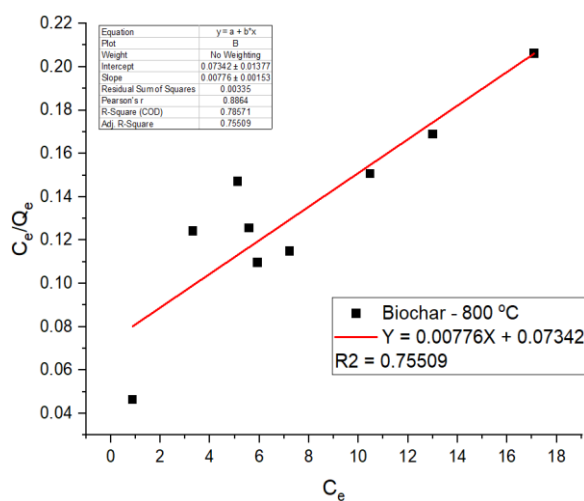


Figure 4. Langmuir and Freundlich adsorption isotherm models of biochar

#### 4. CONCLUSION

This study successfully synthesized biochar material by a simple one-stage calcination method of coffee grounds in anaerobic conditions at 800 °C. Experiments were conducted to evaluate the CV adsorption capacity of biochar in the solution environment. The results showed that the biochar sample could remove the CV crystal violet dye well with high adsorption capacity. CV adsorption kinetics were studied according to the Langmuir and Freundlich isotherm adsorption models. The results showed that the biochar sample was more suitable for the Freundlich isotherm adsorption model than the Langmuir model because the correlation coefficient  $R^2$  value was higher. Therefore, biochar made from coffee grounds is a cheap adsorbent material capable of treating crystal violet dyes, especially colored dyes in textile wastewater.

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## IZVOD

### BIOUGALJ NAPRAVLJEN OD MLEVENE KAFE ZA ADSORPCIJU KRISTALNO LJUBIČASTE BOJE

*Biougalj je pripremljen anaerobnim zagrevanjem mlevene kafe na 800 °C tokom 3 sata. Pripremljeni uzorak je korišćen za adsorbovanje kristalno ljubičaste boje u vodenom okruženju. Rezultat je pokazao da uzorak biougla ima dobar kapacitet adsorpcije kristalno ljubičaste boje, sa maksimalnim kapacitetom adsorpcije od 128,87 (mg/g). Stoga, jeftin biougljen pripremljen od taloga kafe ima veliki potencijal da se koristi kao adsorbujući materijal za uklanjanje kristalno ljubičaste boje u vodenom okruženju.*

**Ključne reči:** biougalj, kristalno ljubičasta boja, mlevena kafa, izotermni modeli, kinetika adsorpcije, tretman otpadnih voda

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